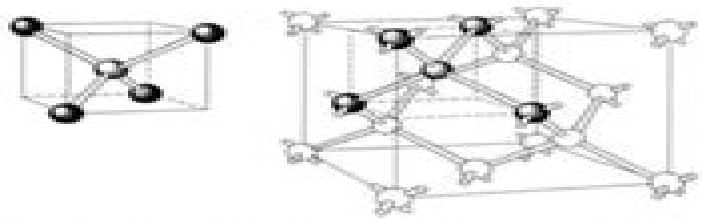


***Fig. Tetrahedral* sp³ bonding (very hard!)**

The diamond structure has an FCC space lattice and a basis comprising two identical atoms. Alternatively, the structure may be viewed as two interpenetrating FCC lattices, one displaced relative to the other along a body diagonal by a quarter of its length.

The atoms on four atoms shown inside the main FCC sub lattice are from other interpenetrated FCC sub lattices.

Relatively few elements exist in the diamond structure. Carbon is the most obvious example, but the importance of this structure in solid state physics is mainly due to the fact that this is the natural state of the semiconductors silicon and germanium.

Many two-element compounds form in a diamond-like structure in which the first element occupies one FCC sublattice and the second element the other. This structure is named after its prime example, ZnS.

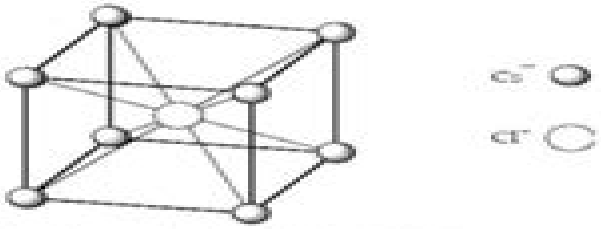
Position of Corbon atoms are at

First kind: (000), (100), (010), (001), (101), (110), (011), (111), (½,0, ½) (½,½,0). (1, ½,½), (½,½,1)

Second kind: (¼, ¼, ¼), (¾ , ¾,¾) (¾ , ¾,¼), ((¼, ¼, ¾,)

CsCl

It contains atoms at the corners of a cube and at the body centre of the cube. Note that it is NOT a body centered cubic structure (as the atom at the centre is different to the atom at the corners of the cube).



In the cesium chloride (CsCl) structure, the space lattice is simple cubic. There are two such simple cubic lattices which are separated by one half of the body diagonal of the unit cube. One lattice is occupied by Cs atoms, the other by Cl atoms.

Each atom has 8 nearest neighbors of the opposite kind and is bound to these 8 atoms which are arranged at the corners of the surrounding hexahedron (cube). The bonding is typically ionic. The dimensions of the CsCl unit cell are $a = b = c = 4.123$ Å

Problems:

Obtain the miller indices of a plane which intercepts at a, b/2, 3c in a simple cubic unit cell.

Solution $x = a$, $y = b/2$ and $z = 3c$

Engineering Physics 1 Year Notes Crystal Structures

Jacqueline K. Barrett

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World Directory of Crystallographers Allan L. Bednowitz,Armin P. Segmüller,2013-04-17 A brief historical account of the background leading to the publication of the first four editions of the World Directory of Crystallographers was presented by G Boom in his preface to the Fourth Edition published late in 1971 That edition was produced by traditional typesetting methods from compilations of biographical data prepared by national Sub Editors The major effort required to produce a directory by manual methods provided the impetus to use computer techniques for the Fifth Edition The account of the production of the first computer assisted Directory was described by S C Abrahams in the preface of the Fifth Edition Computer composition which required a machine readable data base offered several major advantages The choice of typeface and range of characters was flexible Corrections and additions to the data base were rapid and once established it was hoped updating for future editions would be simple and inexpensive The data base was put to other Union uses such as preparation of mailing labels and formulation of lists of crystallographers with specified common fields of interest The Fifth Edition of the World Directory of Crystallographers was published in June of 1977 the Sixth in May of 1981 The Subject Indexes for the Fifth and Sixth Editions were printed in 1978 and 1981 respectively both having a limited distribution **World Directory of Crystallographers** ,2013-11-11 **Crystals and the Crystal Structure of Metallic Compounds** C. G. James,1938

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Engineering Properties of Ceramics James F. Lynch,C. G. Ruderer,Battelle Memorial Institute,W. H. Duckworth,1966 This Databook of engineering properties of ceramics supersedes the Materials Selection Handbook issued in October 1963 as RTD TDR 63 4102 and its two supplements dated May 1964 and April 1965 The materials covered are principally refractory inorganic compounds in the form of monophase crystalline ceramics Exceptions are the coverage of selected refractory

composite materials and intermetallic compounds as well as metalloid elements of boron silicon and germanium Excluded are glasses carbons graphites and compounds melting below about 2000 F 1200 C Properties given are those pertinent to structural engineering The data are in a format designed to permit comparison and selection of materials as well as to provide details on each specific material Discussions of factors affecting property values and structural application of brittle materials also are included *Scientific Personnel Bulletin* United States. Office of Naval Research,1946 **Glaciological Notes** ,1973

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