



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or covalent.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - CO_2
 - CaCl_2
 - H_2S
 - H_2SO_4
- Write the formula for the following compounds:
 - Carbon tetrachloride
 - Carbon disulfide
 - Carbon monoxide
 - Diphosphorus pentoxide
 - Diphosphorus trioxide
 - Diphosphorus pentasulfide
 - Diphosphorus pentasulfide
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SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride. $\text{H}-\text{F}$
- Draw the Lewis structure for carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Draw the Lewis structure for methane. C
- Draw the Lewis structure for ethane. $\text{H}-\text{C}-\text{C}-\text{H}$
- Draw the Lewis structure for ethyne. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw the Lewis structure for acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$



SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by sp^3 ? The diagram shows the sp^3 hybrid orbitals.
- Which orbital and sub-orbital are represented by sp^2 ? The diagram shows the sp^2 hybrid orbitals.

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Anthony S. Fauci



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Table of Contents Covalent Bonding Pearson Answers

1. Understanding the eBook Covalent Bonding Pearson Answers
 - The Rise of Digital Reading Covalent Bonding Pearson Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Pearson Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonding Pearson Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Pearson Answers
 - Personalized Recommendations
 - Covalent Bonding Pearson Answers User Reviews and Ratings
 - Covalent Bonding Pearson Answers and Bestseller Lists
5. Accessing Covalent Bonding Pearson Answers Free and Paid eBooks
 - Covalent Bonding Pearson Answers Public Domain eBooks
 - Covalent Bonding Pearson Answers eBook Subscription Services
 - Covalent Bonding Pearson Answers Budget-Friendly Options
6. Navigating Covalent Bonding Pearson Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Bonding Pearson Answers Compatibility with Devices
 - Covalent Bonding Pearson Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonding Pearson Answers
 - Highlighting and Note-Taking Covalent Bonding Pearson Answers
 - Interactive Elements Covalent Bonding Pearson Answers
8. Staying Engaged with Covalent Bonding Pearson Answers

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonding Pearson Answers
9. Balancing eBooks and Physical Books Covalent Bonding Pearson Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonding Pearson Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Covalent Bonding Pearson Answers
- Setting Reading Goals Covalent Bonding Pearson Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Covalent Bonding Pearson Answers
- Fact-Checking eBook Content of Covalent Bonding Pearson Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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