

Bond Energy

What makes a reaction endothermic or exothermic?

Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** (ΔH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol _{rxn})		Reaction	Change in Energy (kJ/mol _{rxn})
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

1. Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
2. Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
3. Circle the correct word to complete each sentence below.
 - a. When bonds are (broken/formed) there is a positive energy change.
 - b. Breaking bonds is (endothermic/exothermic).
 - c. When bonds are (broken/formed) there is a negative energy change.
 - d. Forming bonds is (endothermic/exothermic).
4. Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
 - a. How do the changes in energy for the reverse reactions compare?
The changes in energy for these reactions are the same, but opposite in sign.
 - b. Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.
If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.

Bond Energy Pogil

Yijin Wang



Bond Energy Pogil:

Reviewing **Bond Energy Pogil**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Bond Energy Pogil**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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Table of Contents Bond Energy Pogil

1. Understanding the eBook Bond Energy Pogil
 - The Rise of Digital Reading Bond Energy Pogil
 - Advantages of eBooks Over Traditional Books
2. Identifying Bond Energy Pogil
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Bond Energy Pogil
 - User-Friendly Interface
4. Exploring eBook Recommendations from Bond Energy Pogil
 - Personalized Recommendations
 - Bond Energy Pogil User Reviews and Ratings

- Bond Energy Pogil and Bestseller Lists
- 5. Accessing Bond Energy Pogil Free and Paid eBooks
 - Bond Energy Pogil Public Domain eBooks
 - Bond Energy Pogil eBook Subscription Services
 - Bond Energy Pogil Budget-Friendly Options
- 6. Navigating Bond Energy Pogil eBook Formats
 - ePub, PDF, MOBI, and More
 - Bond Energy Pogil Compatibility with Devices
 - Bond Energy Pogil Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Bond Energy Pogil
 - Highlighting and Note-Taking Bond Energy Pogil
 - Interactive Elements Bond Energy Pogil
- 8. Staying Engaged with Bond Energy Pogil
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Bond Energy Pogil
- 9. Balancing eBooks and Physical Books Bond Energy Pogil
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Bond Energy Pogil
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Bond Energy Pogil
 - Setting Reading Goals Bond Energy Pogil
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Bond Energy Pogil
 - Fact-Checking eBook Content of Bond Energy Pogil
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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