

Quantum Mechanics – 1st Year Engineering Notes

1. Introduction to Quantum Mechanics

Classical mechanics failed to explain phenomena such as black body radiation, photoelectric effect, and atomic stability. Quantum Mechanics was developed to describe microscopic systems (atoms, electrons, photons).

2. Wave-Particle Duality

Louis de Broglie proposed that every moving particle has an associated wave given by $\lambda = h/p = h/mv$. This was experimentally confirmed by the Davisson-Germer experiment, where electrons showed diffraction patterns when scattered from a nickel crystal.

3. Heisenberg's Uncertainty Principle

It is impossible to simultaneously determine the exact position and momentum of a particle. $\Delta x \Delta p \geq h/4\pi$ and $\Delta E \Delta t \geq h/4\pi$. This implies fundamental limits to precision in quantum systems.

4. Wave Function and Probability

The wave function ψ represents the state of a quantum system. It is generally complex. The probability of finding a particle in a region is given by $|\psi|^2$. Normalization condition: $\int |\psi|^2 dx = 1$.

5. Schrödinger Equation

Time-dependent form: $i\hbar(\partial\psi/\partial t) = -(\hbar^2/2m)(\partial^2\psi/\partial x^2) + V\psi$. Time-independent form: $-(\hbar^2/2m)(\partial^2\psi/\partial x^2) + V\psi = E\psi$. It governs the behavior of quantum particles.

6. Particle in a One-Dimensional Box

A particle is confined between $x=0$ and $x=L$ with infinite potential walls. Boundary conditions: $\psi(0)=0$, $\psi(L)=0$. Solution: $\psi_n(x) = \sqrt{2/L} \sin(n\pi x/L)$. Energy levels: $E_n = n^2\hbar^2/(8mL^2)$, $n = 1, 2, 3, \dots$. Energy is quantized; lowest energy ($n=1$) is not zero (Zero Point Energy).

7. Quantum Tunneling

A quantum particle can penetrate potential barriers higher than its energy, due to the probabilistic nature of wavefunctions. Applications: tunnel diode, alpha decay, scanning tunneling microscope (STM).

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Brendan G. Carr



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