Stoichiometry Worksheet #1

Mole-to-Mole Problems

1) $K_3PO_4 + Al(NO_3)_3 \rightarrow 3\ KNO_3 + AlPO_4$

How many moles of potassium nitrate are produced when **7.70 moles of aluminum nitrate react** with excess potassium phosphate?

$$\frac{7.70 \text{ mol } Al(NO_3)_3 \quad 3 \text{ mol } KNO_3}{1 \text{ mol } Al(NO_3)_3} = \boxed{23.1 \text{ mol } KNO_3}$$

2) $N_2 + 3\ H_2 \rightarrow 2\ NH_3$

How many moles of hydrogen are needed to completely react with 2.60 moles of nitrogen?

$$\frac{2.60 \text{ molN2} \quad 3\text{molH2}}{1\text{Mol N2}} = \boxed{7.80 \text{ mol N2}}$$

3) $N_2 + 3\ H_2 \rightarrow 2\ NH_3$

How many moles of ammonia are produced with 11.30 moles of nitrogen?

$$\frac{11.30 \text{ mol N2} \quad 2 \text{ mol NH3}}{1 \text{ mol N2}} = \boxed{22.6 \text{ mol NH3}}$$

4) $C_3H_8 + 5\ O_2 \rightarrow 3\ CO_2 + 4\ H_2O$

How many moles of oxygen are necessary to react completely with 4.0 moles of propane?

$$\frac{4.0 \text{ mol C3H8} \quad 5\ O2}{1 \text{ mol C3H8}} = \boxed{20\ O2}$$

Mole-to-Mass and Mass-to-Mole Problems:

5) $2\ AgNO_3 + BaCl_2 \rightarrow 2\ AgCl + Ba(NO_3)_2$

How many grams of barium chloride are needed to react completely with 5.00 moles of silver nitrate?

$$\frac{5.00 \text{ mol } AgNO_3 \quad 1 \text{ mol } BaCl_2}{2 \text{ mol } AgNO_3} \cdot \frac{208.23 \text{ g } BaCl_2}{1 \text{ mol } BaCl_2} = \boxed{521 \text{ g } BaCl_2}$$

Solution Stoichiometry Worksheet Answers

Arturo Cuomo

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Table of Contents Solution Stoichiometry Worksheet Answers

1. Understanding the eBook Solution Stoichiometry Worksheet Answers
    - The Rise of Digital Reading Solution Stoichiometry Worksheet Answers
    - Advantages of eBooks Over Traditional Books
2. Identifying Solution Stoichiometry Worksheet Answers
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an Solution Stoichiometry Worksheet Answers
    - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Stoichiometry Worksheet Answers
    - Personalized Recommendations
    - Solution Stoichiometry Worksheet Answers User Reviews and Ratings

- Solution Stoichiometry Worksheet Answers and Bestseller Lists
5. Accessing Solution Stoichiometry Worksheet Answers Free and Paid eBooks
    - Solution Stoichiometry Worksheet Answers Public Domain eBooks
    - Solution Stoichiometry Worksheet Answers eBook Subscription Services
    - Solution Stoichiometry Worksheet Answers Budget-Friendly Options
6. Navigating Solution Stoichiometry Worksheet Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Solution Stoichiometry Worksheet Answers Compatibility with Devices
    - Solution Stoichiometry Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Solution Stoichiometry Worksheet Answers
    - Highlighting and Note-Taking Solution Stoichiometry Worksheet Answers
    - Interactive Elements Solution Stoichiometry Worksheet Answers
8. Staying Engaged with Solution Stoichiometry Worksheet Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Solution Stoichiometry Worksheet Answers
9. Balancing eBooks and Physical Books Solution Stoichiometry Worksheet Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Solution Stoichiometry Worksheet Answers
10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
11. Cultivating a Reading Routine Solution Stoichiometry Worksheet Answers
    - Setting Reading Goals Solution Stoichiometry Worksheet Answers
    - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solution Stoichiometry Worksheet Answers
    - Fact-Checking eBook Content of Solution Stoichiometry Worksheet Answers
    - Distinguishing Credible Sources

13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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