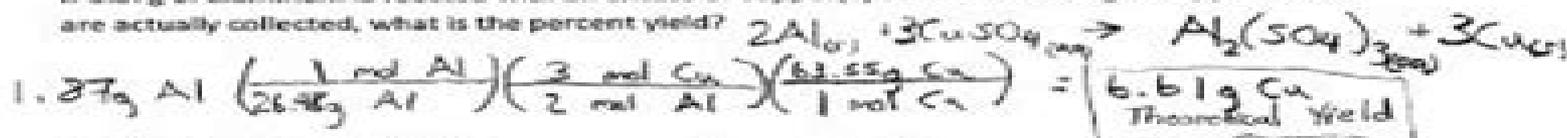


Stoichiometry Worksheet #2

1. If 1.87g of aluminum is reacted with an excess of copper(II) sulfate and 3.74g of copper metal are actually collected, what is the percent yield?



$$\frac{3.74g \text{ Cu (Actual Yield)}}{6.61 \text{ (Theoretical)}} \times 100 = \boxed{56.6\% \text{ Yield}}$$

2. 0.531 mol of iron(III) oxide are reacted with an excess of carbon monoxide and 57.8g of iron is produced. What is the percent yield? Hint: The other product of the reaction is carbon dioxide.



$$\frac{57.8g \text{ (Actual)}}{59.3g \text{ (Theor)}} \times 100 = \boxed{97.5\% \text{ Yield}}$$

3. Calcium carbonate can be decomposed by heating to produce solid calcium oxide and carbon dioxide gas. If you wanted to produce 5.55L of carbon dioxide gas, what mass of calcium carbonate would have to be used? Assume 100% yield and conditions at STP.



$$CaCO_3 \rightarrow \begin{matrix} 40.08 \\ 12.01 \\ 3(16.00) \end{matrix}$$

$$5.55 \text{ L } CO_2 \left(\frac{1 \text{ mol } CO_2}{22.4 \text{ L } CO_2} \right) \left(\frac{1 \text{ mol } CaCO_3}{1 \text{ mol } CO_2} \right) \left(\frac{100.09g \text{ } CaCO_3}{1 \text{ mol } CaCO_3} \right) = \boxed{24.8g \text{ } CaCO_3}$$

4. Aluminum chloride, an inexpensive reagent used in many industrial processes, is made by reacting scrap aluminum with chlorine. If 2.70g of aluminum are reacted with 4.05g of chlorine, which is the limiting reagent and how many grams of the excess reagent remain when the reaction is complete?



$$4.05g \text{ Cl}_2 \left(\frac{1 \text{ mol } Cl_2}{70.90g \text{ Cl}_2} \right) \left(\frac{2 \text{ mol } AlCl_3}{3 \text{ mol } Cl_2} \right) = .037 \text{ mol } AlCl_3$$

Actual Product Cl_2 is limiting

$$.100 \text{ mol Al} - .037 \text{ mol Al} = .063 \text{ mol Al (Product)}$$

$$.063 \text{ mol Al} \left(\frac{26.98g \text{ Al}}{1 \text{ mol Al}} \right) = 1.67g \text{ of Al in Excess}$$

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Rachel S Tattersall



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