

4. The connection between the temperature and pressure of a gas in the process just considered (the so-called Joule-Thomson process) is best expressed in terms of a thermodynamic derivative  $j \equiv (\partial T / \partial P)_H$ . Show that

$$j \equiv \left( \frac{\partial T}{\partial P} \right)_H = \frac{V}{C_P} (T\alpha - 1)$$

where  $\alpha$  is the thermal expansion coefficient, and  $C_P$  is the heat capacity at constant pressure.

5. For the media described by the equation of state given in 2., , investigate how the sign of  $j$  depends on the parameters  $A$  and  $n_0$  and state variables. State the conditions for  $j < 0$ ,  $j = 0$ , and  $j > 0$ .



# Thermal Expansion Problems And Solutions

**SB Merriam**



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