

AP Chemistry - Chemical Kinetics

Name: _____ Date: _____ Period: _____

- After five half-life periods for a first order reaction, what fraction of reactant remains?
- If you plot $1/[\text{reactant}]$ versus time and observe a straight line, what is the order of the reaction? If $\ln[\text{reactant}]$ is plotted versus time, and a straight line of negative slope is observed, what is the order of the reaction?
- Indicate whether each of the following statements is true or false. Change the wording of each false statement to make it true.
 - In stepwise reactions the rate-determining step is the slow one.
 - It is possible to change the rate constant for a reaction by changing the temperature.
 - As a first order reaction proceeds at a constant temperature, the rate remains constant.
 - The rate constant for a reaction is independent of reactant concentrations.
- For each reaction below give the relative rates of formation of products and disappearance of reactants using coefficients and rate change symbols.
 - $2\text{O}_2(\text{g}) \rightarrow 3\text{O}_2(\text{g})$
 - $2\text{HOF}(\text{g}) \rightarrow 2\text{HF}(\text{g}) + \text{O}_2(\text{g})$
 - $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$
- The reaction $\text{CO}(\text{g}) + \text{NO}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{NO}(\text{g})$ is second order in NO_2 and zero order in CO at temperatures less than 500K.
 - Write the rate law for the reaction.
 - How will the reaction rate change if the NO_2 concentration is halved?
- For a reaction of the type $\text{A} + \text{B} \rightarrow \text{C}$, it is experimentally observed that doubling the concentration of B causes the reaction rate to be increased fourfold, but doubling the concentration of A has no effect on the rate. Determine the initial rate equation.
- A study of the reaction $2\text{A} + \text{B} \rightarrow \text{C} + \text{D}$ gave the following results:

Exprmnt #	Init Con (M)		Init Rate (M/s)
	[A] ₀	[B] ₀	
1.	0.10	0.05	6.0×10^{-3}
2.	0.20	0.05	1.2×10^{-2}
3.	0.30	0.05	1.8×10^{-2}
4.	0.20	0.15	1.1×10^{-1}

- What is the rate law for this reaction?
- What is the overall order of the reaction?
- Could this reaction occur in a single step?

- The decomposition of N_2O_5 (to give NO_2 and O_2) follows the rate law "rate = $k[\text{N}_2\text{O}_5]^n$ " where k is 5.0×10^{-4} per second at a particular temperature.
 - What is the half-life of this reaction?
 - How long does it take for the N_2O_5 concentration to drop to one tenth of its original value?

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web feb 10 2015 i ran across this page yesterday and was fascinated by the inner workings of the power split device psd i also noticed that there was enough information there for me to put together an animation and working interactive demo of the psd with controls for input from ice and mg2 so i spent a couple hours putting it together in maya

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web nov 6 2016 introduction many resources and threads exist that describe the prius power split device they correctly describe how rpms can be calculated and how torque is split but i did not find any that would correctly describe how the power is split the solution seems easy at first power is rpm multiplied by torque

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