

The de Broglie Wavelength

- ✧ A particle of large mass and ordinary speed has too small a wavelength to be detected by conventional means.
- ✧ A tiny particle—such as an electron—moving at typical speed has a detectable wavelength.

$$\begin{aligned}\lambda &= \frac{h}{p} = \frac{h}{mv} = \frac{6.63 \times 10^{-34} \text{ J}\cdot\text{s}}{(9.11 \text{ kg})(9.11 \times 10^{-31} \text{ kg})(1 \times 10^6 \text{ m/s})} = 6.63 \times 10^{-35} \text{ m} \\ &= 7.29 \times 10^{-10} \text{ m}\end{aligned}$$

Conceptual Physics Chapter 38 Answer

Zhe Xu



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