

For a first order reaction, $t_{1/2} = \frac{0.6932}{k}$

$$k = \frac{0.6932}{6.932 \times 10^4}$$

$$k = 10^{-5} \text{ s}^{-1}$$

$$k = \left(\frac{1}{t} \right) \ln \left(\frac{[A_0]}{[A]} \right)$$

$$10^{-5} \text{ s}^{-1} \times 100 \times 60 \text{ s} = \ln \left(\frac{[A_0]}{[A]} \right)$$

$$0.06 = \ln \left(\frac{[A_0]}{[A]} \right)$$

$$\frac{[A_0]}{[A]} = e^{0.06}$$

$$\frac{[A_0]}{[A]} = 1.06$$

$$\therefore \frac{[A_0] - [A]}{[A_0]} \times 100 \%$$

$$= \left(1 - \frac{[A]}{[A_0]} \right) \times 100 \%$$

$$= \left(1 - \frac{1}{1.06} \right) \times 100 \%$$

$$= 5.6 \%$$

Chemical Kinetics Problems And Solutions

JS Bruner



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