

## practice test ch 7 & 8

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. Which of the following best describes the rates of chemical reaction?
- most chemical reactions occur very slowly
  - most chemical reactions occur at moderate rates
  - most chemical reactions occur very rapidly
  - chemical reactions have a wide range of rates, from extremely fast to extremely slow
- \_\_\_\_\_ 2. Which of the following is true of the rates of most chemical reactions?
- the initial rate is cannot be measured
  - the initial rate is faster than the rate later in time
  - the initial rate is the same as the rate later in time
  - the initial rate is slower than the rate later in time
- \_\_\_\_\_ 3. Which of the following is true of effective collisions?
- the number of effective collisions decreases as the temperature is increased
  - the number of effective collisions determines the reaction rate
  - both a. and b.
  - Neither a. nor b.
- \_\_\_\_\_ 4. Which of the following is true of effective collisions?
- the number of effective collisions determines the reaction rate
  - the number of effective collisions increases as the temperature is increased
  - both a. and b.
  - Neither a. nor b.
- \_\_\_\_\_ 5. Which of the following is true?
- decreasing the activation increases the reaction rate
  - the activation energy for an exothermic reaction is negative
  - both a. and b.
  - Neither a. nor b.
- \_\_\_\_\_ 6. In an energy diagram for a chemical reaction what species appears at the highest energy?
- the catalyst
  - the products
  - the reactants
  - the transition state
- \_\_\_\_\_ 7. Given that the reaction  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$  is exothermic, which of the following is true of the reaction  $2 \text{H}_2\text{O}(\text{g}) \rightarrow 2 \text{H}_2(\text{g}) + \text{O}_2(\text{g})$ ?
- its activation energy is lower than that of  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
  - its activation energy is the same as that of  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
  - its activation energy is higher than that of c. its activation energy is the same as that of  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
  - there is no relationship between its activation energy and that of  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
- \_\_\_\_\_ 8. In an energy diagram for an endothermic chemical reaction which of the following is true?
- the energy of the products is lower than that of the reactants
  - the energy of the transition state is higher than that of the reactants
  - both a. and b.
  - Neither a. nor b.
- \_\_\_\_\_ 9. Consider a reaction such as  $\text{A}_2(\text{g}) + \text{B}_2(\text{g}) \rightarrow 2\text{AB}(\text{g})$ . What can we say about the mechanism of the reaction?
- if the reaction is fast it occurs in a single step
  - if the reaction is fast it occurs in multiple steps
  - if the reaction is slow it occurs in multiple steps

# Modern Chemistry Chapter 7 Test Answer Key

**Rosina Ehmann**



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### **Table of Contents Modern Chemistry Chapter 7 Test Answer Key**

1. Understanding the eBook Modern Chemistry Chapter 7 Test Answer Key
  - The Rise of Digital Reading Modern Chemistry Chapter 7 Test Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 7 Test Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Chapter 7 Test Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter 7 Test Answer Key
  - Personalized Recommendations
  - Modern Chemistry Chapter 7 Test Answer Key User Reviews and Ratings
  - Modern Chemistry Chapter 7 Test Answer Key and Bestseller Lists
5. Accessing Modern Chemistry Chapter 7 Test Answer Key Free and Paid eBooks
  - Modern Chemistry Chapter 7 Test Answer Key Public Domain eBooks
  - Modern Chemistry Chapter 7 Test Answer Key eBook Subscription Services

- Modern Chemistry Chapter 7 Test Answer Key Budget-Friendly Options
- 6. Navigating Modern Chemistry Chapter 7 Test Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Modern Chemistry Chapter 7 Test Answer Key Compatibility with Devices
  - Modern Chemistry Chapter 7 Test Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 7 Test Answer Key
  - Highlighting and Note-Taking Modern Chemistry Chapter 7 Test Answer Key
  - Interactive Elements Modern Chemistry Chapter 7 Test Answer Key
- 8. Staying Engaged with Modern Chemistry Chapter 7 Test Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Chapter 7 Test Answer Key
- 9. Balancing eBooks and Physical Books Modern Chemistry Chapter 7 Test Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Chapter 7 Test Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Chapter 7 Test Answer Key
  - Setting Reading Goals Modern Chemistry Chapter 7 Test Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Chapter 7 Test Answer Key
  - Fact-Checking eBook Content of Modern Chemistry Chapter 7 Test Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
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