

4. A student adds 15.0 mL of water to a 25.0 mL sample of 0.80 M  $\text{Li}_3\text{PO}_4$  solution.

$$FC = IC \times \frac{V}{FV}$$

- a) Calculate the final  $[\text{Li}^+]$  (2 marks)

$$\text{Final } [\text{Li}_3\text{PO}_4] = 0.80 \text{ M} \times \frac{25.0 \text{ mL}}{40.0 \text{ mL}} = 0.50 \text{ M}$$



$$\text{final } [\text{Li}^+] = 1.5 \text{ M}$$

- b) Calculate the final  $[\text{PO}_4^{3-}]$  (1 mark)



$$\text{final } [\text{PO}_4^{3-}] = 0.50 \text{ M}$$

5. Calculate the mass of solid  $\text{K}_2\text{CO}_3$  which must be added to 500.0 mL of water in order to make a solution in which the  $[\text{K}^+]$  is 3.0 M. (3 marks)



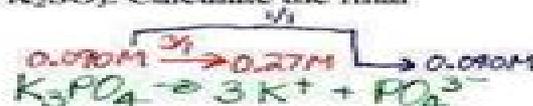
$$\text{mol} = \text{M} \times \text{L} = 1.50 \text{ M} \times 0.5000 \text{ L} = 0.750 \text{ mol } \text{K}_2\text{CO}_3$$

$$\text{mass} = 0.750 \text{ mol } \text{K}_2\text{CO}_3 \times \frac{138.2 \text{ g } \text{K}_2\text{CO}_3}{1 \text{ mol } \text{K}_2\text{CO}_3} = 103.65 \text{ g}$$

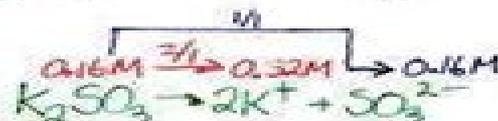
$$\text{mass } \text{K}_2\text{CO}_3 = 104 \text{ g}$$

6. 60.0 mL of 0.45 M  $\text{K}_3\text{PO}_4$  is mixed with 240.0 mL of 0.20 M  $\text{K}_2\text{SO}_4$ . Calculate the final concentration of all three ions in the solution. (3 marks)

$$\text{final } [\text{K}_3\text{PO}_4] = 0.45 \text{ M} \times \frac{60.0 \text{ mL}}{300.0 \text{ mL}} = 0.090 \text{ M}$$



$$\text{final } [\text{K}_2\text{SO}_4] = 0.20 \text{ M} \times \frac{240.0 \text{ mL}}{300.0 \text{ mL}} = 0.16 \text{ M}$$



$$\text{final } [\text{K}^+] = 0.27 \text{ M} + 0.32 \text{ M} = 0.59 \text{ M}$$

$$\text{Answer: } [\text{PO}_4^{3-}] = 0.090 \text{ M} \quad [\text{SO}_4^{2-}] = 0.16 \text{ M} \quad [\text{K}^+] = 0.59 \text{ M}$$

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**CL Gary**



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## **Table of Contents Acid Base Titration Lab Answer Key**

1. Understanding the eBook Acid Base Titration Lab Answer Key
  - The Rise of Digital Reading Acid Base Titration Lab Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Titration Lab Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acid Base Titration Lab Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Titration Lab Answer Key
  - Personalized Recommendations
  - Acid Base Titration Lab Answer Key User Reviews and Ratings
  - Acid Base Titration Lab Answer Key and Bestseller Lists
5. Accessing Acid Base Titration Lab Answer Key Free and Paid eBooks
  - Acid Base Titration Lab Answer Key Public Domain eBooks
  - Acid Base Titration Lab Answer Key eBook Subscription Services
  - Acid Base Titration Lab Answer Key Budget-Friendly Options

6. Navigating Acid Base Titration Lab Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Acid Base Titration Lab Answer Key Compatibility with Devices
  - Acid Base Titration Lab Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acid Base Titration Lab Answer Key
  - Highlighting and Note-Taking Acid Base Titration Lab Answer Key
  - Interactive Elements Acid Base Titration Lab Answer Key
8. Staying Engaged with Acid Base Titration Lab Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acid Base Titration Lab Answer Key
9. Balancing eBooks and Physical Books Acid Base Titration Lab Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acid Base Titration Lab Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Acid Base Titration Lab Answer Key
  - Setting Reading Goals Acid Base Titration Lab Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acid Base Titration Lab Answer Key
  - Fact-Checking eBook Content of Acid Base Titration Lab Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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