

Date Performed: 10/26/09

Date Submitted: 11/02/09

Instructor:

Name:

Partner:

Acid-Base Titrations

Objective

- To standardize a sodium hydroxide solution (NaOH) with a hydrochloric acid solution (HCl) of known concentration by the reaction.
- To determine the molar mass of an unknown, monoprotic acid
- To determine the percentage of Acetylsalicylic Acid in a tablet of aspirin.

Experimental Data

Part A-Standardization of Sodium Hydroxide Solution

Titration	Vol NaOH (mL)
1	20.00
2	20.21

Part B-Molar Mass of an Unknown Monoprotic Acid

Sample	Mass Beaker (g)	Mass Sample + Beaker (g)	Vol NaOH (mL)
1	119.50	120.50	20.95
2	88.32	89.32	19.95

Part C-Percentage of Acetylsalicylic Acid in a Commercial Aspirin Tablet

Mass Beaker (g)	Mass Tablet +Beaker (g)	Vol NaOH (mL)
122.08	122.46	8.49

Sample Calculations

1) Molarity NaOH

$(\text{Molarity HCl})(\text{Volume HCl}) = (\text{Molarity NaOH})(\text{Volume NaOH})$

$(0.200\text{M})(25.00\text{mL}) = (\text{Molarity NaOH})(20.00\text{mL})$

$5.00 = X(20.00)$

Trail 1: Molarity of NaOH = 0.250M NaOH

Trail 2: Molarity of NaOH = 0.247M NaOH

Average = 0.249M NaOH

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Table of Contents Acid Base Titrations Pre Lab Answers Chem Fax

1. Understanding the eBook Acid Base Titrations Pre Lab Answers Chem Fax
 - The Rise of Digital Reading Acid Base Titrations Pre Lab Answers Chem Fax
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Titrations Pre Lab Answers Chem Fax
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Base Titrations Pre Lab Answers Chem Fax
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Titrations Pre Lab Answers Chem Fax
 - Personalized Recommendations
 - Acid Base Titrations Pre Lab Answers Chem Fax User Reviews and Ratings
 - Acid Base Titrations Pre Lab Answers Chem Fax and Bestseller Lists
5. Accessing Acid Base Titrations Pre Lab Answers Chem Fax Free and Paid eBooks
 - Acid Base Titrations Pre Lab Answers Chem Fax Public Domain eBooks
 - Acid Base Titrations Pre Lab Answers Chem Fax eBook Subscription Services

- Acid Base Titrations Pre Lab Answers Chem Fax Budget-Friendly Options
- 6. Navigating Acid Base Titrations Pre Lab Answers Chem Fax eBook Formats
 - ePub, PDF, MOBI, and More
 - Acid Base Titrations Pre Lab Answers Chem Fax Compatibility with Devices
 - Acid Base Titrations Pre Lab Answers Chem Fax Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid Base Titrations Pre Lab Answers Chem Fax
 - Highlighting and Note-Taking Acid Base Titrations Pre Lab Answers Chem Fax
 - Interactive Elements Acid Base Titrations Pre Lab Answers Chem Fax
- 8. Staying Engaged with Acid Base Titrations Pre Lab Answers Chem Fax
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Base Titrations Pre Lab Answers Chem Fax
- 9. Balancing eBooks and Physical Books Acid Base Titrations Pre Lab Answers Chem Fax
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Base Titrations Pre Lab Answers Chem Fax
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acid Base Titrations Pre Lab Answers Chem Fax
 - Setting Reading Goals Acid Base Titrations Pre Lab Answers Chem Fax
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acid Base Titrations Pre Lab Answers Chem Fax
 - Fact-Checking eBook Content of Acid Base Titrations Pre Lab Answers Chem Fax
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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