

1. List the intermolecular forces that can be found between molecules and define them.

**Hydrogen bonding**—attraction between the H bound to either F, O, or N, and a lone pair on one of those atoms.

**Dipole-Dipole**—attraction between the  $-$  and  $+$  ends of two molecules.

**London Forces**—small attraction that happens due to the presence of electrons.

2. For the following molecules, draw the Lewis structure, state whether it is polar/nonpolar, and list the intermolecular forces found in the molecules. Which would you expect to have the strongest intermolecular forces?

Water has the strongest intermolecular forces because it has the most hydrogen bonding, dipole-dipole, and London forces.

- a.  $\text{H}_2\text{O}$   
polar, H-bonding, Dip-dip, London



- b.  $\text{CH}_4$



Nonpolar, London

- c.  $\text{CHCl}_3$   
polar, Dip-Dip, London



- d.  $\text{OF}_2$   
polar, Dip-Dip, London



- e.  $\text{NH}_3$   
Polar, H-bonding, Dip-Dip, London



- f.  $\text{PH}_3$



Polar, Dip-Dip, London

# Molecular Geometry And Intermolecular Forces Answer Key

**Michael Mosher, Paul Kelter**



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