

**B. Titration of the Acid of unknown concentration with the Standardized NaOH**

1. Record your unknown ID.
2. Obtain less than 10mL of an unknown acid in a clean, dry 50-mL beaker.
3. Rinse the 10.00-mL pipet with three 2-mL portions of your unknown solution. Obtain ~ 40mL of unknown acid and pipet 10.00mL of it into the rinsed 250-mL flask.
4. Repeat the titration procedure above using the acid of unknown concentration in place of the standard acid.

Set up a data table to organize your data for the H<sub>2</sub>SO<sub>4</sub>/NaOH titration.

	Trial 1	Trial 2	Trial 3
Unknown ID			
V <sub>H<sub>2</sub>SO<sub>4</sub></sub> (mL)	10.00	10.00	10.00
Avg M <sub>NaOH</sub> (M), from Part A			
V <sub>initial, NaOH</sub> (mL)			
V <sub>final, NaOH</sub> (mL)			
Δ V <sub>NaOH</sub> (mL)			
M <sub>H<sub>2</sub>SO<sub>4</sub></sub> (M)			

**Calculations**

In these titrations, one mole of HCl reacts with one mole of NaOH, and one mole of H<sub>2</sub>SO<sub>4</sub> reacts with two moles of NaOH. The molarity, M, of the acid or the base is expressed in moles per liter.

In part A the molarity of the HCl solution is known. The molarity of NaOH may be calculated from the results of part A through the following stoichiometry reflecting the overall equation:



1. M<sub>NaOH</sub>, the molarity of the NaOH calculated from the data. Show one calculation of the NaOH molarity in detail
2. Average value of the three determinations of NaOH molarity.
3. Percent range of the three determinations:

$$\% \text{ range} = \frac{|\text{largest value} - \text{smallest value}|}{\text{average value}} \times 100\%$$

In part B the molarity of the NaOH solution determined from part A is used to compute the molar concentration of the H<sub>2</sub>SO<sub>4</sub> in the solution of unknown concentration:



1. M<sub>H<sub>2</sub>SO<sub>4</sub></sub>, the molarity of the H<sub>2</sub>SO<sub>4</sub> calculated from the data. Show one calculation of the H<sub>2</sub>SO<sub>4</sub> molarity in detail
2. Average value of the three determinations of H<sub>2</sub>SO<sub>4</sub> molarity.
3. Percent range of the three determinations:

$$\% \text{ range} = \frac{|\text{largest value} - \text{smallest value}|}{\text{average value}} \times 100\%$$

# Acid Base Titrations Pre Lab Answers

**Xiang Xie**



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