

$$Z = \frac{(\bar{X} - \mu_{\bar{X}})}{\sigma_{\bar{X}}} = \frac{(\bar{X} - \mu)}{\frac{\sigma}{\sqrt{n}}}$$

where:  $\bar{X}$  = sample mean  
 $\mu$  = population mean  
 $\sigma$  = population standard deviation  
 $n$  = sample size

## BUSS1020 QUANTITATIVE BUSINESS ANALYSIS NOTES

### EXCEL

- Average: =average(first box then press control, shift, arrow down to highlight all, then data box then press control, shift, arrow down to highlight all)
- Average of only some of the people: =averageif(first box then press control, shift, arrow down to highlight all, then press the specific thing you are averaging, then data box then press control, shift, arrow down to highlight all)
- Count: =count(first box then press control, shift, arrow down to highlight all, then data box then press control, shift, arrow down to highlight all)
- =average()
- =mode()
- =median()
- =geomean()
  - If you use Geomean() for rates, you need to first convert the rates to factors, and then convert the answer back to a rate
    - Rate = 0.1
    - Factor = 1.1 (e.g if increasing by 10%)
- =skew
- =kurt
- =quartile.inc()
  - To find quartile 1 = =quartile.inc( , 1)
- =var.s()
- =stdev.s()
- Range =max( )-min( )
- Variance: =var.s( )
- Five number summary
  - quartile.inc( ,here put in the box of 0,1,2,3,4 is)
  - Then drag it down to the other boxes
- Z score
  - =(B2-\$mean)/\$standard deviation
  - Drag down onto all the boxes
- Outliers
  - =COUNTIF( , "<-3")+COUNTIF( , ">3")
- Two way table
  - Highlight all data and press pivot table
- Binomial = binom.dist(box, \$n, \$pi, false)
- Poisson = poisson.dist(success, mean, false)
- Hypergeometric = hypgeom.dist(x, \$sample size (n), \$events of interest (x), \$population size (N), false)
- Normal distribution
  - = norm.dist(x, mean, SD, true/false)

# Business Quantitative Analysis Notes

**DP Hallahan**



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web for example take the reaction of lead ii nitrate with hydrochloric acid to form lead ii chloride and nitric acid shown below  
$$\text{Pb(NO}_3)_2 \text{ (aq)} + 2 \text{HCl (aq)} \rightarrow \text{PbCl}_2 \text{ (s)} + 2 \text{HNO}_3$$

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$$\text{Molecular: AgNO}_3 \text{ (aq)} + \text{KCl (aq)} \rightarrow \text{AgCl (s)} + \text{KNO}_3 \text{ (aq)}$$
  
$$\text{Total Ionic: Ag}^+ \text{ (aq)} + \text{NO}_3^- \text{ (aq)} + \text{K}^+ \text{ (aq)} + \text{Cl}^- \text{ (aq)} \rightarrow \text{AgCl (s)} + \text{K}^+ \text{ (aq)} + \text{NO}_3^- \text{ (aq)}$$

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