

Specific Heat:

Chem. Worksheet 10-1

Name: _____

Use the data in the table for answers. Use 3 significant figures.

Substance	Specific Heat Capacity (J/g °C)
water	4.184
aluminum	0.897
iron	0.449
glass	0.840
gold	0.129
lead	0.128

Initial Temperature	
aluminum	10.0 °C (20 °F)
water	20.0 °C (68 °F)
glass	20.0 °C (68 °F)
lead	20.0 °C (68 °F)

- How much heat energy is required to raise a quantity of water of mass 100.0 g by 10.0 °C by heating? **4184 J**
- An amount of water (100.0 g) at 20.0 °C is mixed with an amount of lead (100.0 g) at 100.0 °C. Assuming that no heat is lost to the surroundings, what is the final temperature of the mixture? **26.2 °C**
- Assuming the temperature of the water is 20.0 °C, how much heat energy is required to raise 100.0 g of gold? **840 J**
- What heat is required to raise a quantity of lead of mass 100.0 g from 20.0 °C to 100.0 °C? **1290 J**
- An amount of lead of mass 100.0 g is heated to 100.0 °C. How much heat energy is required to raise the temperature of the lead to 20.0 °C? **-1290 J**
- An amount of 100.0 g of lead is mixed with a quantity of glass of mass 100.0 g. The temperature of the lead is 100.0 °C. Assuming that no heat is lost to the surroundings, what is the final temperature of the mixture? **26.2 °C**
- What is the amount of energy that must be added to a quantity of 100.0 g of water at 20.0 °C to raise its temperature to 100.0 °C? **4184 J**
- How much energy is required to raise a quantity of 100.0 g of water from 20.0 °C to 100.0 °C? **4184 J**
- What heat energy is required to raise a quantity of 100.0 g of water from 20.0 °C to 100.0 °C? **4184 J**
- What heat energy is required to raise a quantity of 100.0 g of water from 20.0 °C to 100.0 °C? **4184 J**
- An amount of 100.0 g of water is heated from 20.0 °C to 100.0 °C. How much heat energy is required to raise the temperature of the water to 100.0 °C? **4184 J**
- An amount of 100.0 g of water is heated from 20.0 °C to 100.0 °C. How much heat energy is required to raise the temperature of the water to 100.0 °C? **4184 J**

4184 J **4184 J**

Specific Heat Chem Worksheet 16 1 Answer Key

**Thirumalaisamy P.
Velavancorresponding**

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