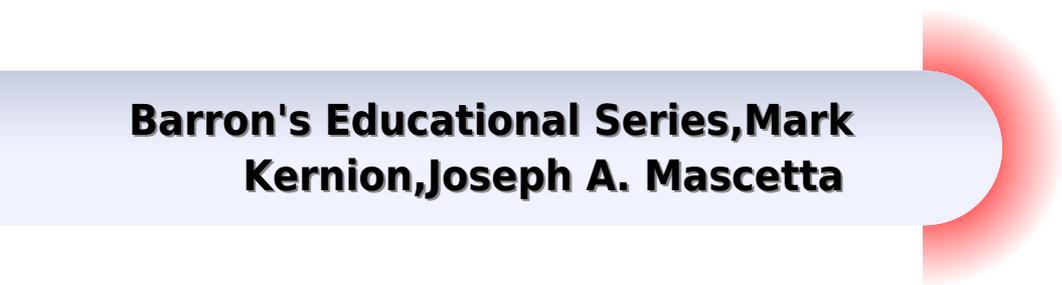


**CHAPTER 7 REVIEW***Chemical Formulas and Chemical Compounds***SECTION 7-1****SHORT ANSWER** Answer the following questions in the space provided.

1.   c   In a Stock name such as iron(III) sulfate, the roman numeral tells us \_\_\_\_\_
- how many atoms of Fe are in one formula unit
  - how many sulfate ions can be attached to the iron atom
  - the charge on each Fe ion
  - the total positive charge of the formula unit
2.   c   The result of changing a subscript in a correctly written chemical formula is to \_\_\_\_\_
- change the number of moles represented by the formula
  - change the charges on the other ions in the compound
  - change the formula so that it no longer represents the compound it previously represented
  - have no effect on the formula
3. The explosive TNT has the molecular formula  $C_7H_5(NO_2)_3$ .
- 4 elements   a. How many elements make up this compound?
- 6 oxygen atoms   b. How many oxygen atoms are present in one molecule of  $C_7H_5(NO_2)_3$ ?
- 21 atoms   c. How many atoms in total are present in one molecule of  $C_7H_5(NO_2)_3$ ?
- $4.2 \times 10^{24}$  atoms   d. How many atoms are present in a sample of  $2 \times 10^{23}$  molecules of  $C_7H_5(NO_2)_3$ ?
4. How many atoms are present in each of these formula units?
- 11 atoms   a.  $Ca(HCO_3)_2$
- 45 atoms   b.  $C_{12}H_{22}O_{11}$
- 10 atoms   c.  $Fe(ClO_2)_3$
- 9 atoms   d.  $Fe(ClO_3)_2$
5.    $N_2O_5$    a. What is the formula for the compound dinitrogen pentoxide?
- carbon(IV) sulfide   b. What is the Stock name for the covalent compound  $CS_2$ ?
- $H_2SO_3$    c. What is the formula for sulfurous acid?
- phosphoric acid   d. What is the name for the acid  $H_3PO_4$ ?

# Chapter 9 Chemical Names And Formulas Practice Problems Answers

**Barron's Educational Series, Mark  
Kernion, Joseph A. Mascetta**



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*Understanding Chemistry* Fred M. Dewey, 1994      **General Organic and Biological Chemistry** Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features      *Addison-Wesley Chemistry*, 1990      *Student's Guide to Introduction to Chemical Principles* by Edward I. Peters, 2d Ed Peter P. Berlow, 1978      *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students      Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998      **Scientific Soapmaking** Kevin M. Dunn, 2010 Scientific Soapmaking

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**Partial Solutions Manual** Darrell D. Ebbing, David Bookin, 1998

**Prentice Hall Chemistry**, 2000 *New Technical Books* New York Public Library, 1925

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*Chemical Engineering Catalog*, 1925 *Chemistry* James E. Brady, Fred Senese, 2004-02-04 Publisher Description

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*The Software Encyclopedia 2001*, 2001

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