

Name _____
AP Chapter 8 Bonding



In this activity we will learn how to draw Lewis structures, calculate formal charge and examine resonance. We will also observe the Octet rule and exceptions to this rule.

AP Chem. Lewis Dot Structure Worksheet

Write Lewis Dot Structures and calculate the formal charges for the following compounds:

- | | |
|------------------------------|-------------------------------------|
| 1. SO_2 | 18. H_2O_2 |
| 2. CO_3^{2-} | 19. C_2H_6 |
| 3. NH_3 | 20. CH_3OH |
| 4. CH_4 | 21. C_2H_4 |
| 5. H_2 | 22. C_4H_{10} |
| 6. HF | 23. $\text{C}_2\text{H}_5\text{OH}$ |
| 7. H_2O | 24. ClO^- |
| 8. N_2 | 25. H_2S_2 |
| 9. HCl | 26. HNO_3 |
| 10. H_2S | 27. H_2SO_4 |
| 11. PF_3 | 28. CH_3NO_2 |
| 12. CH_3COOH | 29. C_2H_2 |
| 13. NH_4^+ | 30. ClO_2^- |
| 14. NO_2^- | 31. PO_4^{3-} |
| 15. SO_4^{2-} | 32. CO |
| 16. NO_2^+ | 33. CO_2 |
| 17. OH^- | |

Lewis Dot Structure Practice Problems With Answers

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



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Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter,1993

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