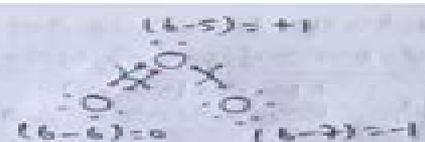


$0 \rightarrow 1 \rightarrow O_2$



→ Stability of same molecules of diff sp

- ① More stable due to symmetrical charge distribution.
- ② More stable due to absence of charge.
- ③ More stable due to -ve charge present on high EN element.

Avg. formal charge & Avg. X-O bond order in oxo-anions:-

$$\text{Avg. f.c.} = \frac{\text{total charge present on O-atom}}{\text{No. of O-atom where -ve charge can be placed}}$$

$$\text{Avg. X-O bond order} = \frac{\text{total no. of X-O bonds}}{\text{no. of X-O pairs participate in resonance}}$$



F.C on 'O' → -2/3
X-O bond order → 4/3

-2/4
6/4

TRICK	FC
= O →	0
- O →	-1
≡ O or	+1
= O -	

Covalent bond

↳ Formed by the equal sharing of e⁻

* Covalency:- Total no. of covalent bond (σ & π) formed by an atom in a molecule.

VALENCY (combining capacity)

① Covalency → (Total no. of σ & π bonds)

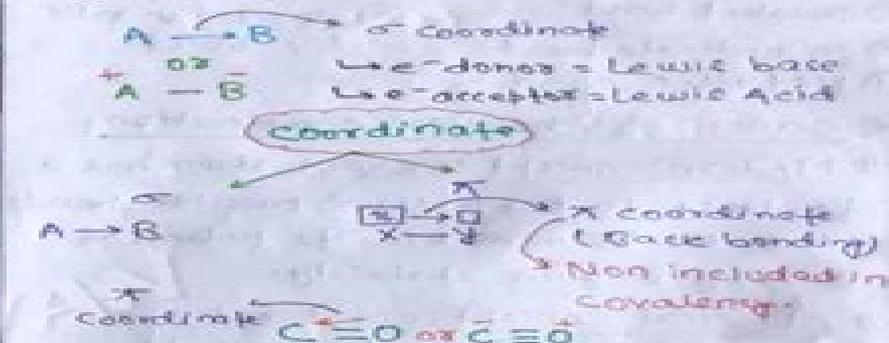
② Electronegativity → Combining capacity in ionic comp. E.g. $Na^+Cl^- \rightarrow Na^+ = +1$
 $Cl^- = -1$

NOTE:-

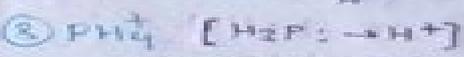
- ① Max covalency of 2nd period element can not be more than 4 due to absence of d-orbitals.
- ② 3rd period elements can extend their covalency beyond 4 by using vacant 3d orbitals.
- ③ 'F' can show only 1 covalency.

Coordinate bond

↳ Formed by the unequal sharing of e⁻
↳ formed by b/w e⁻ donor & e⁻ acceptor



Species in which σ coordinate bond (dative bond) is present:



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