

## 1. Basic Notation of Light

- 1. Light is the transfer of energy.
- 2. The main properties of light include frequency and wavelength. Frequency is generally denoted the wavelength by the Greek letter  $\lambda$  or  $\lambda_{light}$  where  $\lambda$  is wavelength,  $\nu$  is frequency, and  $c$  is speed of light  $3 \times 10^8$  m/s.
- 3. Light has two properties: electromagnetic waves and particles. It is called light in collective terms.
- 4. The wavelength is denoted by  $\lambda$  and the frequency is denoted by  $\nu$  or  $\nu_{light}$  where  $\lambda$  is wavelength of the photon,  $\nu$  is Planck's constant multiplied by  $\lambda$ , and  $\nu$  is frequency.
- 5. The quantum properties of electromagnetic systems are described by equations:  $E = h\nu$  where  $E$  is energy,  $h$  is Planck's constant, and  $\nu$  is frequency or wavelength.
- 6. Energy

$E = h\nu$  where  $\nu$  is frequency or  $\lambda$

$E = hc/\lambda$  where  $\lambda$  is wavelength,  $h$  is Planck's constant, and  $c$  is speed of light

$\lambda = hc/E$  where  $E$  is energy,  $h$  is Planck's constant, and  $c$  is speed of light

$\nu = E/h$

$\nu = E/h$

$\lambda = hc/E$  where  $E$  is energy,  $h$  is Planck's constant, and  $c$  is speed of light

### Common Questions

1. How does the wavelength of light change as frequency changes? Does the energy of photons also change?
  - a. As frequency increases, wavelength decreases. The energy of photons also increases.
2. A photon's energy is directly proportional to its frequency. How does the wavelength of the light change?
  - a. As frequency increases, wavelength decreases.
3. The energy of a photon is directly proportional to its frequency. How does the wavelength of the light change?
  - a. As frequency increases, wavelength decreases.
4. A photon's energy is directly proportional to its frequency. How does the wavelength of the light change?
  - a. As frequency increases, wavelength decreases.
5. In what way is the energy of a photon related to its wavelength?
  - a. Energy is directly proportional to frequency, which is inversely proportional to wavelength.
6. Calculate the energy of a photon with a wavelength of 400 nm.
  - a.  $E = hc/\lambda = (6.626 \times 10^{-34} \text{ J}\cdot\text{s})(3 \times 10^8 \text{ m/s}) / (400 \times 10^{-9} \text{ m}) = 4.97 \times 10^{-19} \text{ J}$
  - b.  $E = hc/\lambda = (6.626 \times 10^{-34} \text{ J}\cdot\text{s})(3 \times 10^8 \text{ m/s}) / (400 \times 10^{-9} \text{ m}) = 4.97 \times 10^{-19} \text{ J}$
7. Calculate the energy of a photon with a wavelength of 700 nm.
  - a.  $E = hc/\lambda = (6.626 \times 10^{-34} \text{ J}\cdot\text{s})(3 \times 10^8 \text{ m/s}) / (700 \times 10^{-9} \text{ m}) = 2.84 \times 10^{-19} \text{ J}$
  - b.  $E = hc/\lambda = (6.626 \times 10^{-34} \text{ J}\cdot\text{s})(3 \times 10^8 \text{ m/s}) / (700 \times 10^{-9} \text{ m}) = 2.84 \times 10^{-19} \text{ J}$
8. How does the energy of a photon change as its wavelength changes?
  - a. Energy is inversely proportional to wavelength.
9. How does the energy of a photon change as its frequency changes?
  - a. Energy is directly proportional to frequency.
10. How does the energy of a photon change as its wavelength changes?
  - a. Energy is inversely proportional to wavelength.
11. How does the energy of a photon change as its frequency changes?
  - a. Energy is directly proportional to frequency.
12. Calculate the energy of a photon with a wavelength of 400 nm.
  - a.  $E = hc/\lambda = (6.626 \times 10^{-34} \text{ J}\cdot\text{s})(3 \times 10^8 \text{ m/s}) / (400 \times 10^{-9} \text{ m}) = 4.97 \times 10^{-19} \text{ J}$
  - b.  $E = hc/\lambda = (6.626 \times 10^{-34} \text{ J}\cdot\text{s})(3 \times 10^8 \text{ m/s}) / (400 \times 10^{-9} \text{ m}) = 4.97 \times 10^{-19} \text{ J}$

# Chemquest 11 Answers

**R Pring**



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