



# CHEMISTRY

PERIODIC TABLE OF THE ELEMENTS

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
H	He	Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra	Ac	Rf	Mn	Co	Ni	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe		

### PERIODIC TRENDS

- Atomic Size**
  - Increases from top-left to bottom-right.
  - Increases from top to bottom in a group.
  - Decreases from left to right in a period.
- Ionization Potential**
  - Increases from bottom-left to top-right.
  - Increases from left to right in a period.
  - Decreases from top to bottom in a group.
- Electronegativity**
  - Increases from bottom-left to top-right.
  - Increases from left to right in a period.
  - Decreases from top to bottom in a group.

### ATOMIC STRUCTURE AND THE PERIODIC TABLE

- Atomic Spectrum**
  - Discrete lines.
  - Wavelength increases from violet to red.
  - Frequency decreases from violet to red.
- Photoelectric Effect**
  - Electrons are emitted only if the frequency of light is above a certain threshold.
  - Maximum kinetic energy of electrons is directly proportional to the frequency of light.

$$E = hf$$

where  $E = 2.0 \times 10^{-18} \text{ J}$ ,  $h = 6.626 \times 10^{-34} \text{ J s}$ , and  $f = ?$

### IONIZATION ENERGY AND ATOMIC SIZE



Ionization energy is the energy required to remove an electron from an atom in its ground state.

- Electronic Structure of Atoms**
  - Electrons are arranged in shells.
  - Maximum number of electrons in a shell is  $2n^2$ .
  - Outermost shell is the valence shell.



### THE WAVE MECHANICAL MODEL

- Electrons are described as waves.
- Electrons are confined to orbitals.
- Orbitals are regions of space where the probability of finding an electron is high.

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- Electrons are confined to orbitals.
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### PERIODIC TABLE

- The periodic table is a table of elements.
- Elements are arranged in groups and periods.
- Groups are vertical columns.
- Periods are horizontal rows.
- Properties of elements vary across the table.
- Metals are on the left, non-metals on the right.
- Alkali metals are in Group 1.
- Halogens are in Group 17.
- Noble gases are in Group 18.
- Transition metals are in Groups 3-10.
- Inner transition metals are in Groups 13-18.
- Properties of elements vary across the table.
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### WORKED EXAMPLE

- WORKED EXAMPLE**
  - Calculate the wavelength of light that has a frequency of  $5.0 \times 10^{14} \text{ s}^{-1}$ .
  - Calculate the energy of a photon of light that has a wavelength of  $4.0 \times 10^{-7} \text{ m}$ .

# Chemistry Quickstudy Reference Guides Academic

**Mark Jackson**



## **Chemistry Quickstudy Reference Guides Academic:**

**Chemistry** BarCharts Inc,2012-05-31     Chemistry Mark D. Jackson,BarCharts, Inc,2001-01-24 This 4 page laminated guide is essential for students or individual looking for Chemistry knowledge The QuickStudy Chemistry guide includes information on atomic structure nuclear chemistry atomic quantum numbers orbitals electronic atoms types of matter reactions physical processes nomenclature stoichiometry chemical interactions bonding models gasesthermodynamics theories and much more     *Organic Chemistry Fundamentals* Mark Jackson,2015-12-31 Quick Reference for the core essentials of a subject and class that is challenging at best and that many students struggle with In 6 laminated pages our experienced chemistry author and professor gathered key elements organized and designed to use along with your text and lectures as a review before testing or as a memory companion that keeps key answers always at your fingertips As many students have said a must have study tool Suggested uses o Quick Reference instead of digging into the textbook to find a core answer you need while studying use the guide to reinforce quickly and repeatedly o Memory refreshing your memory repeatedly is a foundation of studying have the core answers handy so you can focus on understanding the concepts o Test Prep no student should be cramming but if you are there is no better tool for that final review     **QuickStudy Academic** Mark Jackson (PhD and Science Professor), BarCharts best selling quick reference to chemistry has been updated and expanded in this new edition With updated content and an additional panel of information this popular guide is not only an essential companion for students in introductory chemistry courses but also a must have refresher for students in higher level courses Author Mark D Jackson PhD a scientist and university chemistry professor has a gift for making the complicated subject of chemistry interesting and easy to understand without the fluff In this new edition you will find more coverage of the subject helpful illustrations chemical problems and practical applications making this a study tool you won t want to be without     *Organic Chemistry Reactions* BarCharts, Inc,Mark Jackson,2015-12-31 Quick Reference for the core essentials of a subject and class that is challenging at best and that many students struggle with In 6 laminated pages our experienced chemistry author and professor gathered key elements organized and designed to use along with your text and lectures as a review before testing or as a memory companion that keeps key answers always at your fingertips As many students have said a must have study tool Suggested uses o Quick Reference instead of digging into the textbook to find a core answer you need while studying use the guide to reinforce quickly and repeatedly o Memory refreshing your memory repeatedly is a foundation of studying have the core answers handy so you can focus on understanding the concepts o Test Prep no student should be cramming but if you are there is no better tool for that final review     **Chemistry Equations and Answers** Mark Jackson,2006-02-22 This 6 page study guide contains basic chemistry analysis and concepts designed specifically to aid science students     Chem Lab Basics Mark Jackson, Frank Miskevich,2018-05 With the study of molecules and how they interact with one another theory is important but the lab is where the true excitement of chemistry lies This 6 page laminated

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