

## AP Chemistry Chapter 4 Practice Exam

- D 1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.



How much sulfur (in grams) is produced from 21.5 grams of  $\text{O}_2$ ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$21.5 \text{ g O}_2 \left| \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right| \frac{2 \text{ SO}_2}{3 \text{ O}_2} \left| \frac{3 \text{ S}}{\text{SO}_2} \right| \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

- B 2. Iron is produced from its ore by the reactions:



How many moles of  $\text{O}_2(g)$  are needed to produce 4.6 moles of  $\text{Fe}(s)$ ?

A) 2.3 mole  $\text{O}_2$

B) 3.5 mole  $\text{O}_2$

C) 4.6 mole  $\text{O}_2$

D) 6.9 mole  $\text{O}_2$

E) 13.8 mole  $\text{O}_2$

$$4.6 \text{ mol Fe}(s) \left| \frac{3 \text{ CO}}{2 \text{ Fe}} \right| \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

- B 3. Suppose the reaction  $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$  is carried out starting with 153 g of  $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of  $\text{H}_2\text{SO}_4$ . How much phosphoric acid will be produced?

A) 76.7 g

B) 51.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$153 \text{ g Ca}_3\text{PO}_4)_2 \left| \frac{1 \text{ mol Ca}_3\text{PO}_4)_2}{312 \text{ g Ca}_3\text{PO}_4)_2} \right| \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3\text{PO}_4)_2} =$$

$$76.8 \text{ g H}_2\text{SO}_4 \left| \frac{1 \text{ mol H}_2\text{SO}_4}{98 \text{ g H}_2\text{SO}_4} \right| \frac{2 \text{ H}_3\text{PO}_4}{3 \text{ H}_2\text{SO}_4} =$$

- B 4.  $\text{SO}_2$  reacts with  $\text{H}_2\text{S}$  as follows:



When 7.50 g of  $\text{H}_2\text{S}$  reacts with 12.75 g of  $\text{SO}_2$ , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.

C) 0.0216 moles of  $\text{H}_2\text{S}$  remain.

D) 1.13 g of  $\text{H}_2\text{S}$  remain.

E)  $\text{SO}_2$  is the limiting reagent.

$$7.50 \text{ g H}_2\text{S} \left| \frac{1 \text{ mol H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \right| \frac{3 \text{ S}}{2 \text{ H}_2\text{S}} \left| \frac{32 \text{ g S}}{1 \text{ mol S}} \right| = 10.89$$

$$12.75 \text{ g SO}_2 \left| \frac{1 \text{ mol SO}_2}{64.06 \text{ g SO}_2} \right| \frac{3 \text{ S}}{1 \text{ SO}_2} \left| \frac{32 \text{ g S}}{1 \text{ mol S}} \right| = 19.13$$

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