Stoichiometry

SECTION 1

Multiple Choice Write the letter of the correct answer on the line provided.

1. ____a____ When a sample of a compound is decomposed, the mass of the compound is equal to the mass of the elements that make up the compound.
   (a) conservation
   (b) composition
   (c) multiple proportions
   (d) law of conservation of mass
   (e) law of definite proportions
   (f) law of multiple proportions
   (g) law of conservation of energy
   (h) law of conservation of matter
   (i) law of conservation of mass and energy
   (j) law of conservation of mass and energy
   (k) law of conservation of mass and energy
   (l) law of conservation of mass and energy
   (m) law of conservation of mass and energy
   (n) law of conservation of mass and energy
   (o) law of conservation of mass and energy
   (p) law of conservation of mass and energy
   (q) law of conservation of mass and energy
   (r) law of conservation of mass and energy
   (s) law of conservation of mass and energy
   (t) law of conservation of mass and energy
   (u) law of conservation of mass and energy
   (v) law of conservation of mass and energy
   (w) law of conservation of mass and energy
   (x) law of conservation of mass and energy
   (y) law of conservation of mass and energy
   (z) law of conservation of mass and energy

2. ____a____ Which of the following statements are true about the law of conservation of mass? Select all that apply.
   (a) The mass of a closed system remains constant over time.
   (b) The mass of a closed system increases over time.
   (c) The mass of a closed system decreases over time.
   (d) The mass of a closed system increases or decreases over time.
   (e) The mass of a closed system increases or decreases over time.
   (f) The mass of a closed system increases or decreases over time.
   (g) The mass of a closed system increases or decreases over time.
   (h) The mass of a closed system increases or decreases over time.
   (i) The mass of a closed system increases or decreases over time.
   (j) The mass of a closed system increases or decreases over time.
   (k) The mass of a closed system increases or decreases over time.
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   (t) The mass of a closed system increases or decreases over time.
   (u) The mass of a closed system increases or decreases over time.
   (v) The mass of a closed system increases or decreases over time.
   (w) The mass of a closed system increases or decreases over time.
   (x) The mass of a closed system increases or decreases over time.
   (y) The mass of a closed system increases or decreases over time.
   (z) The mass of a closed system increases or decreases over time.

3. ____a____ Which of the following statements are true about the law of conservation of mass? Select all that apply.
   (a) The mass of a closed system remains constant over time.
   (b) The mass of a closed system increases over time.
   (c) The mass of a closed system decreases over time.
   (d) The mass of a closed system increases or decreases over time.
   (e) The mass of a closed system increases or decreases over time.
   (f) The mass of a closed system increases or decreases over time.
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   (w) The mass of a closed system increases or decreases over time.
   (x) The mass of a closed system increases or decreases over time.
   (y) The mass of a closed system increases or decreases over time.
   (z) The mass of a closed system increases or decreases over time.

4. Consider the following equation. Select all that apply.

   $2H_2(g) + O_2(g) \rightarrow 2H_2O(g)$

   a. Balance the equation.

   $__H_2(g) + __O_2(g) \longrightarrow __H_2O(g)$

   Assume that 0.5 mol H_2 is present in the reaction mixture with 0.25 mol O_2.

   b. How many moles of H_2O are produced? _____ c. Which reactant is the limiting reagent? _____ d. How many moles of the excess reagent are left over? _____

   e. How many moles of H_2O are produced? _____ f. How many moles of O_2 are left over? _____

   g. How many moles of H_2O are produced? _____ h. How many moles of O_2 are left over? _____

Chapter 9 Review Stoichiometry Section 3

Hussin A.Rothana

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