

## Chapter 1 – Kinetics - Practice Test #1

Name - \_\_\_\_\_

The answers to this test are posted online

1. Consider the following reaction mechanism:
- step 1:  $M + X \rightarrow MX$
- step 2:  $MX + A \rightarrow D + X$
- The chemical species MX is a(n)
- A. catalyst  
B. inhibitor **B**  
C. final product  
D. reaction intermediate
2. Consider the following reaction:  $2N_2O_{5(g)} \rightarrow 4NO_{2(g)} + O_{2(g)}$
- At a certain temperature the rate of decomposition of  $2N_2O_{5(g)}$  is  $2.5 \times 10^{-6}$  mol/s. The rate of formation of  $NO_2$  is
- A.  $1.0 \times 10^{-5}$  mol/s  
B.  $1.3 \times 10^{-6}$  mol/s  
C.  $2.5 \times 10^{-6}$  mol/s  
D.  $5.0 \times 10^{-6}$  mol/s
3. Which of the following factors affect the rates of both homogeneous and heterogeneous reactions:
- I nature of the reactants  
II presence of a catalyst  
III temperature of system  
IV concentration of reactants
- A. I and IV only  
B. II and III only  
C. II, III, and IV only  
D. I, II, III, and IV
4. Which of the following equations represents an endothermic reaction?
- A.  $N_2O_{5(g)} + 59 \text{ kJ} \rightarrow 2NO_{2(g)}$   
B.  $2H_{2(g)} + O_{2(g)} \rightarrow 2H_2O_{(l)} + 572 \text{ kJ}$   
C.  $2BrCl_{(g)} - 29.3 \text{ kJ} \rightarrow Br_{2(g)} + Cl_{2(g)}$   
D.  $2H_{2(g)} + O_{2(g)} \rightarrow 2H_2O_{(l)} \quad \Delta H = -572 \text{ kJ}$
5. Consider the potential energy diagram. The activation energy for the reverse reaction is:
- A. 30 kJ  
B. 140 kJ  
C. 170 kJ  
D. 200 kJ
- 
- The diagram shows a potential energy curve. The vertical axis is labeled 'PE (kJ)' and the horizontal axis is labeled 'Progress of the reaction'. The curve starts at a reactant energy level, rises to a peak, and then falls to a product energy level. The reactant energy level is 170 kJ higher than the product energy level. The activation energy for the reverse reaction is 30 kJ.

# Chemistry Chemical Kinetics Practice Questions Answers

**Michael Seilmaier**



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