

Chemistry 338 2nd Semester Final Exam Review Sheet

District Assessment

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten	¹⁸⁴ W ⁺² 74	74	74	110	72	184	+2

- Name the compound
 - Fe₂O₃ – iron (III) oxide
 - Cl₂O₈ – dichlorine octaoxide
- Write the formula
 - Trisulfur Dinitride - S₃N₂
 - Zinc Hydroxide – Zn(OH)₂
 - Tin (IV) oxide - SnO₂
- The pressure of a gas is 843 mm Hg at 24°C. What will the pressure be at standard temperature? Volume is constant.

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \frac{843 \text{ mmHg}}{297 \text{ K}} = \frac{P_2}{273 \text{ K}} \quad P_2 = 775 \text{ mm Hg}$$
- The volume of a gas is 81.7 mL when the pressure is 119 kPa and the temperature is 73°C. What is the volume at STP?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \frac{(119 \text{ kPa})(81.7 \text{ mL})}{346 \text{ K}} = \frac{(101.3 \text{ kPa})(V_2)}{273 \text{ K}} \quad V_2 = 75.7 \text{ mL}$$
- Between Be and Ba, which has a larger ionization energy? Larger atomic radius?

Larger ionization energy – Be
Larger atomic radius - Ba
- Balance the equation & classify the type of reaction:

$$2\text{Al} + 3 \text{CuSO}_4 \rightarrow 1 \text{Al}_2(\text{SO}_4)_3 + 3 \text{Cu} \quad \text{single replacement}$$
- Predict the products & classify the type of reaction:

$$\text{H}_3\text{PO}_4 + 3 \text{NaOH} \rightarrow 3 \text{H}_2\text{O} + \text{Na}_3\text{PO}_4 \quad \text{double replacement}$$
- Calculate the Molar Mass: (NH₄)₂CO₃
96 g/mole
- Convert 2.75 moles of N₂O₅ to molecules

$$\frac{2.75 \text{ moles}}{1} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}} = 1.66 \times 10^{24} \text{ molecules N}_2\text{O}_5$$

Solutions

Chemistry Spring Semester Final Exam Review Sheet

G Thomas



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