

## CHAPTER 9 REVIEW

**Stoichiometry****SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- \_\_\_\_\_ The coefficients in a chemical equation represent the
  - masses in grams of all reactants and products.
  - relative number of moles of reactants and products.
  - number of atoms of each element in each compound in a reaction.
  - number of valence electrons involved in a reaction.
- \_\_\_\_\_ Which of the following would not be studied within the topic of stoichiometry?
  - the mole ratio of Al to Cl in the compound aluminum chloride
  - the mass of carbon produced when a known mass of sucrose decomposes
  - the number of moles of hydrogen that will react with a known quantity of oxygen
  - the amount of energy required to break the ionic bonds in  $\text{CaF}_2$
- \_\_\_\_\_ A balanced chemical equation allows you to determine the
  - mole ratio of any two substances in the reaction.
  - energy released in the reaction.
  - electron configuration of all elements in the reaction.
  - reaction mechanism involved in the reaction.
- \_\_\_\_\_ The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
  - reaction sequence.
  - bond energy.
  - mole ratio.
  - element proportion.
- Given the reaction represented by the following unbalanced equation:  
 $\text{N}_2\text{O}(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}_2(\text{g})$ 
  - Balance the equation.

\_\_\_\_\_

b. What is the mole ratio of  $\text{NO}_2$  to  $\text{O}_2$ ?

\_\_\_\_\_

c. If 20.0 mol of  $\text{NO}_2$  form, how many moles of  $\text{O}_2$  must have been consumed?

\_\_\_\_\_

d. Twice as many moles of  $\text{NO}_2$  form as moles of  $\text{N}_2\text{O}$  are consumed. True or False?

\_\_\_\_\_

e. Twice as many grams of  $\text{NO}_2$  form as grams of  $\text{N}_2\text{O}$  are consumed. True or False?

# Chapter 9 Stoichiometry Section 2 Worksheet

**Leo J. Malone, Theodore O. Dolter**



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**Handbook on Material and Energy Balance Calculations in Material Processing, Includes CD-ROM** Arthur E. Morris, H. Alan Fine, Gordon Harold Geiger, 2011-09-06 This book approaches the subject of material and energy balances from two directions First it emphasizes the fundamental principles of the conservation of mass and energy and the consequences of these two principles Second it applies the techniques of computational chemistry to materials processing and introduces new software developed by the author especially for material and heat balances The third edition reflects the changes in the professional engineer's practice in the last 30 years reflecting the dramatic shift away from metallurgical engineering and the extractive industry towards materials engineering A large and growing number of recent graduates are employed in such fields as semiconductor processing environmental engineering and the production and processing of advanced and exotic materials for aerospace electronic and structural applications The advance in computing power and software for the desktop computer has significantly changed the way engineers make computations and the biggest change comes from the computational approach used to solve problems The spreadsheet program Excel is used extensively throughout the text as the main computational engine for solving material and energy balance equations and for statistical analysis of data The use of Excel and the introduction of the add in programs enables the study of a range of variables on critical process parameters and emphasis is placed on multi device flowsheets with recycle bypass and purge streams whose material and heat balance equations were previously too complicated to solve by the normally used hand calculator The Excel based program FlowBal helps the user set up material and heat balance equations for processes with multiple streams and units

*Holt Chemistry* Ralph Thomas Myers, 2004

**MCAT Comprehensive Review, 2004** Kaplan, Inc, 2003 The bestselling MCAT prep guide is a must for pre med students preparing for the demanding Medical College Admissions Test MCAT The CD features one full length practice test with detailed explanations plus effective tips and strategies

**MCAT Comprehensive Review**, 2004

**Basic Concepts of Chemistry** Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the

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*The Practice of Chemistry* Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can t do chemistry if they can t do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book s unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POCPowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt) 832KB

**Chemical Problem-solving by Dimensional Analysis** Arnold B. Loebel, 1978

**Instructor's Manual and Test Bank to Accompany Basic Concepts of Chemistry** Malone, 2000-04-18

**Basic Concepts of Chemistry, Study Guide** Leo J. Malone, 2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and

critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry. Designed to help readers master the principles of general chemistry. As a prep book it promotes active involvement with the material. There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills. Updated to include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems.

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*Study Guide for Chemical Principles [by] Steven S. Zumdahl* Steven S. Zumdahl, Paul B. Kelter, 1995. The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text. The new edition of the Study Guide includes many new worked out examples.

Basic Chemical Principles Edward I. Peters, 1988. **CHEMISTRY** SILBERBERG, 2003. **Programmed Problem-solving for First-year Chemistry** Arnold B. Loebel, 1983. Basic Chemistry G. William Daub, William S. Seese, 1996. Written in a style and language that users without science backgrounds can understand. This best selling introduction to the basic principles of chemistry draws on the reader's own experiences through analogies and cartoons to learn difficult concepts. The clear systematic thinking approach to problem solving has also been highly praised by reviewers and users alike. Countdown sections in each chapter consisting of five review questions keyed to previous material provide readers with a basis for material introduced in the new chapter. Study exercises found immediately after new topics are introduced reinforce chapter problem material. You and Chemistry marginal application icon relates chemistry to the real world. End of chapter essays entitled Elements and Compounds relate the applications of specific elements or compounds to the reader's life.

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