

Directions: Read carefully and solve the following problems.

1. The complete combustion of Octane results to the production of Carbon dioxide and water as shown in the unbalanced chemical equation below.



- a. Write the balanced chemical reaction equation.
  - b. How much  $CO_2$  in grams will be produced from burning 30 grams of Octane?
  - c. What mass of Oxygen is needed to completely consume 100 grams of Octane?
2. In a certain experiment, Hydrochloric acid is used to decompose Calcium carbonate. The chemical reaction equation is shown below.



- a. After the experiment, the total mass of  $CaCl_2$  produced is 120g. Assuming that the reaction proceeded to completion, how many moles of  $CaCO_3$  were used in the experiment?
  - b. Based on the amount of  $CaCl_2$  produced in the reaction, how much  $CO_2$  in grams effervesced out of the Calcium carbonate used?
3. Hydrogen fluoride is produced by reacting Calcium fluoride with sulfuric acid. A by-product of the process is Calcium sulfate.

# Chapter 3 Stoichiometry Exercises Answers

**James C. Hill**



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